

Thermodynamics, structure, and dynamics of water confined between hydrophobic platesPradeep Kumar,¹ Sergey V. Buldyrev,^{1,2} Francis W. Starr,³ Nicolas Giovambattista,^{1,4} and H. Eugene Stanley¹¹*Center for Polymer Studies and Department of Physics Boston University, 590 Commonwealth Avenue, Boston, Massachusetts 02215, USA*²*Department of Physics, Yeshiva University, 500 West 185th Street, New York, New York 10033, USA*³*Department of Physics, Wesleyan University, Middletown, Connecticut 06459 USA*⁴*Department of Chemical Engineering, Princeton University, Princeton, New Jersey 08544, USA*

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We perform molecular dynamics simulations of 512 waterlike molecules that interact via the TIP5P potential and are confined between two smooth hydrophobic plates that are separated by 1.10 nm. We find that the anomalous thermodynamic properties of water are shifted to lower temperatures relative to the bulk by ≈ 40 K. The dynamics and structure of the confined water resemble bulk water at higher temperatures, consistent with the shift of thermodynamic anomalies to lower temperature. Because of this T shift, our confined water simulations (down to $T=220$ K) do not reach sufficiently low temperature to observe a liquid-liquid phase transition found for bulk water at $T\approx 215$ K using the TIP5P potential, but we see inflections in P - ρ isotherms at lower temperatures presumably due to the presence of a liquid-liquid critical point. We find that the different crystalline structures that can form for two different separations of the plates, 0.7 and 1.10 nm, have no counterparts in the bulk system, and we discuss the relevance to experiments on confined water.

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I. INTRODUCTION

Despite the numerous accomplishments in water research to date [1–7], the topic continues to be the subject of intense interest. In particular, water confined in nanoscale geometries has garnered much recent attention because of its biological and technological importance [8,9]. Confinement can lead to changes in both structural and dynamical properties caused by the interaction with a surface and/or a truncation of the bulk correlation length. Moreover, these changes depend on whether the interactions of water with the wall particles are hydrophilic or hydrophobic [10]. One of the motivations for studying two different kinds of interactions arises from studies of protein folding, since the folding of a protein is influenced by its hydrophobic and hydrophilic interactions with water [11].

It is not clear exactly how the dynamics of liquids depends on the nature of the confining surfaces. The dynamic behavior may change depending on the surface morphology. Simulations of simple liquids show that the dynamics typically *slow down* near a nonattractive rough surface, whereas the dynamics *speed up* near a nonattractive smooth surface [12]. A slowing down of water dynamics near a hydrophilic surface has been experimentally observed [13]. Water confined in Vycor [14,15] has at least two different dynamical regimes arising from the slow dynamics of water near the surface and fast dynamics of water far away from the surfaces [16–20].

One anomaly hypothesized to occur in supercooled water is the emergence of a phase transition line separating liquid states of different densities. This phenomenon is called a liquid-liquid (LL) phase transition [2,21–26]. A LL transition has been seen in a variety of simulation models of water [21,27,28], but is difficult to observe, experimentally, due to the propensity of ice to nucleate at temperatures where a transition is expected. Nonetheless, indirect evidence of a

transition has been found [2,29–33]. Studies of some simple models of liquid also show a LL phase transition [34–38].

Bulk water simulations using the TIP5P potential [39,40] indicate the presence of a LL phase transition ending in a second critical point at $T\approx 217$ K and $\rho\approx 1.13\pm 0.04\text{g/cm}^3$ [27,41]. A LL phase transition has been suggested based on simulations using the ST2 potential confined between smooth plates [42]. A liquid-to-amorphous transition is seen in simulations using the TIP4P potential [43–45] confined in carbon nanotubes [46]. Recent theoretical work [47] suggests that hydrophobic confinement suppresses the LL transition to lower T . Here we aim to determine how confinement between smooth hydrophobic walls affects the location of the LL critical point as well as the overall thermodynamic, dynamic, and structural properties.

The freezing of water in confined spaces is also interesting. On one hand, recent experimental studies of water confined in carbon nanopores show that water does not crystallize even when the temperature is cooled down to 77 K [13]. On the other hand, computer simulation studies show that models of water can crystallize into different crystalline forms when confined between surfaces [48–52]. For example, monolayer ice was found in simulations using the TIP5P model of water [49]. By applying an electric field along lateral directions (directions perpendicular to the confinement direction) another crystalline structure for three molecular layers of water confined between two silica plates was found [50]. Also, bilayer hexagonal ice was found in simulations using the TIP4P model [48]. In general these simulations predict a variety of polymorphs in confined spaces, but the crystalline structures found have yet to be observed in experiments.

This paper is organized as follows: In Sec. II, we provide details of our simulations and analysis methods. Simulation results for the liquid state are provided in Secs. III–V. The

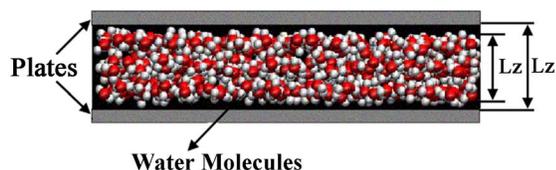


FIG. 1. (Color online) Schematic of the simulated system. Water molecules are confined between two smooth hydrophobic plates in an xy plane, separated by $L_z=1.1$ nm. The figure also indicates L'_z , the effective length of confinement of the water molecules.

crystal states are discussed in Sec. VI, and we conclude with a brief summary in Sec. VII.

II. SIMULATION AND ANALYSIS METHODS

We perform molecular dynamics (MD) simulations of a system composed of $N=512$ waterlike molecules confined between two smooth walls. The molecules interact via the TIP5P pair potential [39], which, like the ST2 [53] potential, treats each water molecule as a tetrahedral, rigid, and non-polarizable unit consisting of five-point sites. Two positive point charges of charge $q_H=0.241e$ (where e is the fundamental unit of charge) are located on each hydrogen atom at a distance 0.09572 nm from the oxygen atom; together they form an HOH angle of 104.52° . Two negative point charges ($q_e=-q_H$) representing the lone pair of electrons (e^-) are located at a distance 0.07 nm from the oxygen atom. These negative point charges are located in a plane perpendicular to the HOH plane and form an e^-Oe^- angle of $\cos^{-1}(1/3) = 109.47^\circ$, the tetrahedral angle. To prevent overlap of molecules, a fifth interaction site is located on the oxygen atom and is represented by a Lennard-Jones (LJ) potential with parameters $\sigma_{OO}=0.312$ nm and $\epsilon_{OO}=0.6694$ kJ/mol.

The TIP5P potential accurately reproduces many water anomalies when no confinement is present [27]. For example, it accurately reproduces the density anomaly at $T=277$ K and $P=1$ atm. Its structural properties compare well with experiments [27,39,40,54]. TIP4P and TIP5P are known to crystallize [27,55] within accessible computer simulation time scales. Moreover, TIP5P shows a “nose-shaped” curve of temperature versus crystallization time [27], a feature found in experimental data on water solutions [56]. TIP5P simulations also show a van der Waals loop in the P - ρ plane at the lowest T accessible with current computation facilities [27]. This loop indicates the presence of a first-order LL transition. Ref. [27] estimates that a LL-transition line ends in a LL critical point C' located at $T_{C'}=217\pm 3$ K, $P_{C'}=340\pm 20$ MPa, and $\rho_{C'}=1.13\pm 0.04$ g/cm³.

In our simulation, water molecules are confined between two infinite smooth planar walls, as shown schematically in Fig. 1. The walls are located at $z=\pm 0.55$ nm, corresponding to a wall-wall separation of 1.1 nm, which results in ≈ 2 – 3 layers of water molecules. Periodic boundary conditions are used in the x and y directions, parallel to the walls.

The interactions between water molecules and the smooth walls are designed to mimic solid paraffin [57] and are given by [58]

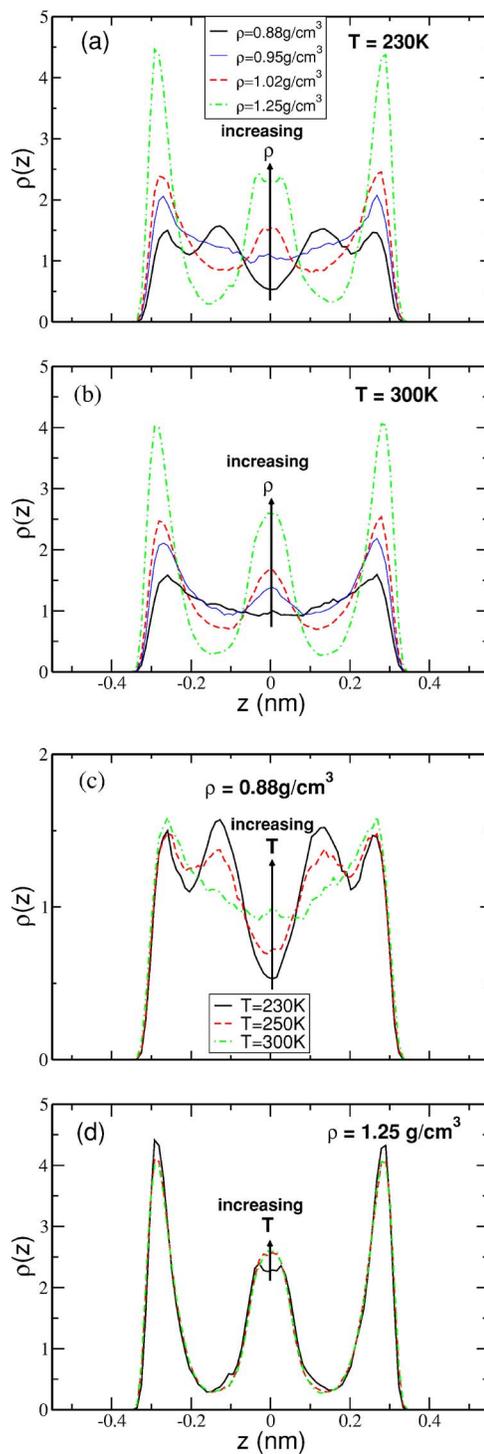


FIG. 2. (Color online) Density of water $\rho(z)$ along the confinement direction for different densities at constant temperatures (a) $T=230$ K and (b) $T=300$ K. (c) Density of water along the confinement direction for different temperatures at constant densities (c) $\rho=0.88$ g/cm³ and (d) $\rho=1.25$ g/cm³. See discussion in text.

$$U(\Delta z) = 4\epsilon_{OW} \left[\left(\frac{\sigma_{OW}}{\Delta z} \right)^9 - \left(\frac{\sigma_{OW}}{\Delta z} \right)^3 \right]. \quad (1)$$

Here Δz is the distance from the oxygen atom of a water molecule to the wall, whereas $\epsilon_{OW}=1.25$ kJ/mole and σ_{OW}

=0.25 nm are potential parameters. The same parameter values were used in previous confined water simulations [57,59].

We perform simulations for 56 state points, corresponding to seven temperatures $T=220, 230, 240, 250, 260, 280,$ and 300 K, and eight densities $\rho=0.80, 0.88, 0.95, 1.02, 1.10, 1.17, 1.25,$ and 1.32 g/cm³ [60]. The range of density values takes into account the fact that the water-wall interactions prevent water molecules from accessing a space near the walls. Our determination of ρ is discussed, in detail, in Sec. III. The raw “geometric” densities used are $\rho=0.60, 0.655, 0.709, 0.764, 0.818, 0.873, 0.927,$ and 0.981 g/cm³.

For each state point, we perform two independent simulations to improve the statistics. We control the temperature using the Berendsen thermostat with a time constant of 5 ps [61] and use a simulation time step of 1 fs, just as in the bulk system [27]. For long-range interactions we use a cutoff of 0.9 nm [39].

We calculate the lateral pressure $P_{\parallel}=(P_{xx}+P_{yy})/2$ using the virial expression for the x and y directions [62]. We obtain the pressure along the transverse direction, P_{\perp} by calculating the total force F_{wall} perpendicular to the wall [42],

$$P_{\perp} = \frac{F_{\text{wall}}}{L_x L_y} = \frac{\sum_{i=1}^N \mathbf{F}_{i,\text{wall}}}{L_x L_y}. \quad (2)$$

Here, $\mathbf{F}_{i,\text{wall}}$ is the force produced by oxygen atom of water molecule i on the wall. Hydrogen atoms do not interact with the wall. In agreement with the simulations of Ref. [59] using the TIP4P model for water and the water-wall interaction given by Eq. (1), we find that the hydrogen atoms of the water molecules near a wall tend to face away from the wall, forming bonds with other molecules. In this sense, the wall interactions are hydrophobic.

III. PROPERTIES OF TIP5P CONFINED WATER

A. Transverse density profile

One of the problems when dealing with liquids in a confined geometry is how to define the density in a consistent way. Using a geometric definition $\rho \equiv Nm/L_x L_y L_z$ (where m is the water molecule mass) underestimates the effective density since the repulsive interactions with the walls prevent molecules from coming too close to the walls. Hence, we want to quantify the effective distance L'_z perpendicular to the walls accessible to the water molecules (Fig. 1), and thus obtain a definition for ρ , which is more readily comparable to the density of a bulk system. To estimate L'_z , we calculate the density profile $\rho(z)$ defined as the geometric density of oxygen centers at z , shown in Fig. 2 for different temperatures and densities. In all cases studied, we observe that molecules cannot access the total available space between the walls, and that the accessible space along the transverse direction does not strongly depend on T and ρ . Hence, we estimate

$$L'_z = L_z - \frac{\sigma_{\text{OW}} + \sigma_{\text{OO}}}{2} = 0.819 \text{ nm} \quad (3)$$

independent of T and ρ [42]; this leads to the effective density

$$\rho \equiv \frac{Nm}{L_x L_y L'_z}. \quad (4)$$

Figure 2(a) shows the effect on $\rho(z)$ of changing ρ at $T=230$ K. Since the typical oxygen-oxygen separation for nearest neighbor in bulk water is 0.28 nm, for the effective wall separation $L'_z=0.819$ nm one would expect that at most three water layers can be accommodated between the walls. At $\rho=1.02$ g/cm³ and $\rho=1.25$ g/cm³, $\rho(z)$ shows three clear maxima indicating the presence of a trilayer liquid. The two maxima next to the walls are the result of water-wall interaction. As density decreases below $\rho=1.02$ g/cm³, the central maximum becomes nearly uniform and, at density $\rho=0.95$ g/cm³, only the two maxima located next to the walls remain. This density corresponds to the least structured liquid. Upon further expansion, the structure of the liquid starts to increase since the bilayer splits into two sublayers for the lowest density $\rho=0.88$ g/cm³. As we will see in Sec. III B, $T=230$ K is below the temperature of maximum density (T_{MD}). Hence, the increase in the structure upon expansion corresponds to the anomalous decrease in entropy upon expansion found in bulk water below the T_{MD} ,

$$\left(\frac{\partial S}{\partial V}\right)_T = \left(\frac{\partial P}{\partial T}\right)_V = -\left(\frac{\partial V}{\partial T}\right)_P \left(\frac{\partial P}{\partial V}\right)_T \leq 0. \quad (5)$$

Figure 2(b) shows the effect on $\rho(z)$ of changing ρ at higher T , $T=300$ K. The densities in Fig. 2(b) are the same as those shown in Fig. 2(a). Similar to findings at $T=230$ K, (i) the liquid at $T=300$ K and $\rho=1.25$ g/cm³ is characterized by a trilayer structure and (ii) reducing the density transforms the trilayer liquid into a bilayer liquid at $\rho=0.88$ g/cm³. However, comparison of Figs. 2(a) and 2(b) shows that, at the lowest ρ studied, the layers of the bilayer structured liquid at $T=300$ K do not split into two sublayers, as is the case of $T=230$ K since $T=300$ lies above the T_{MD} . In fact, we find that at $T=230$ K the sublayers present at $\rho=0.88$ g/cm³ merge into a single layer at $\rho=0.95$ g/cm³, and the resulting $\rho(z)$ resembles that shown in Fig. 2(b) at $\rho=0.88$ g/cm³.

The effect of changing temperature at low density is shown in Fig. 2(c). As discussed above, at low density and low T we observe a bilayer liquid where the two layers split into two sublayers. Increasing T smoothes features of the density profile, namely, (i) the splitting of the sublayers disappears and (ii) the minimum of $\rho(z)$ at $z=0$ nm becomes less pronounced since increasing T increases the entropy. Although a bilayer \rightarrow trilayer liquid crossover is found upon isothermal compression, isochoric heating does not have such an effect.

Figure 2(d) shows the effect of changing T at the high density, $\rho=1.25$ g/cm³. At high ρ , the molecules are not able to displace perpendicular to the walls and the density profile

is almost T independent. Indeed, the liquid has a trilayer structure at all T studied.

B. Low-temperature phase diagram

Next we test how confinement affects the location in the P - T phase diagram of the LL phase transition line and the second critical point found in bulk water simulations using the TIP5P model [27,41]. To determine the phase behavior, we evaluate P as a function of ρ . If there is a LL phase transition, then it should manifest itself in van der Waals “loop” along isothermal paths at low T .

Unlike bulk liquid systems, one must be careful to interpret separately the results involving the lateral pressure P_{\parallel} and those involving the transverse pressure P_{\perp} , since the thermodynamic averages of these quantities will usually be different. Phase separation will only be apparent in P_{\parallel} , since the separation of the plates is too small to allow for the existence of two distinct phases in the transverse direction. In Figs. 3(a) and 3(b), we show P_{\perp} and the lateral pressure P_{\parallel} as functions of density along all seven isotherms.

P_{\perp} is a monotonically increasing function of the density [Fig. 3(a)]. P_{\parallel} is also a monotonic increasing function of ρ , but as T decreases, the isotherms of P_{\parallel} become “flatter” in the region near $\rho \approx 0.95$ g/cm³ [Fig. 3(b)]. The presence of this region is consistent with the possible existence of a van der Waals loop at lower T and, therefore, is consistent with the possible existence of a LL phase transition line ending at a second critical point at a value of T lower than 220 K, the lowest simulated temperature. At 220 K no phase separation occurs, consistent with the simulations of bulk water where the same TIP5P potential gives a second critical point with $T_{C'} = 217 \pm 3$ K [27]. Although we are unable to simulate the temperatures below $T_{C'}$ of the bulk system, comparison of the lowest T isotherm with that of Refs. [27,41] suggests that we will need to go well below the bulk $T_{C'}$ to see phase separation. Thus, our results suggest that the presence of hydrophobic walls shifts a possible second critical point to lower T . Along an isothermal path, the critical point can be located by the point where the slope $(\partial P / \partial \rho)_T$ and curvature $(\partial^2 P / \partial \rho^2)_T$ simultaneously equal zero. We can estimate this point by plotting the value of the minimum slope along each isotherm and extrapolating the slopes to find the T at which the slope is zero. This estimate yields critical temperature $T_{C'} \approx 162 \pm 20$ K (Fig. 4). A crude estimate of the critical pressure $P_{\parallel C}$ shows that $P_{\parallel C}$ smaller than the bulk LL-critical pressure $P_C = 340 \pm 20$ MPa estimated in Ref. [27]. This is consistent with the theoretical predictions in Ref. [47].

Figures 3(c) and 3(d) show the T dependence of P_{\perp} and P_{\parallel} for different densities. $P_{\perp}(T)$ is a monotonic function of T for all ρ . Similar behavior is observed for $P_{\parallel}(T)$ at large ρ . However, for 0.88 g/cm³ $\leq \rho \leq 1.17$ g/cm³, the isochores in the P_{\parallel} - T plane display minima, indicating the presence of a T_{MD} line, defined as the locus of points where $(\partial P / \partial T)_V = 0$ [58]. For $T < T_{MD}$, water confined between hydrophobic walls is anomalous, i.e., it becomes less dense upon cooling. A T_{MD} line has also been found in TIP5P bulk water simulations [27]. Comparison of Fig. 3(c) to Fig. 2(a) of Ref. [27] shows that the T_{MD} locus in confined water shifts to lower T .

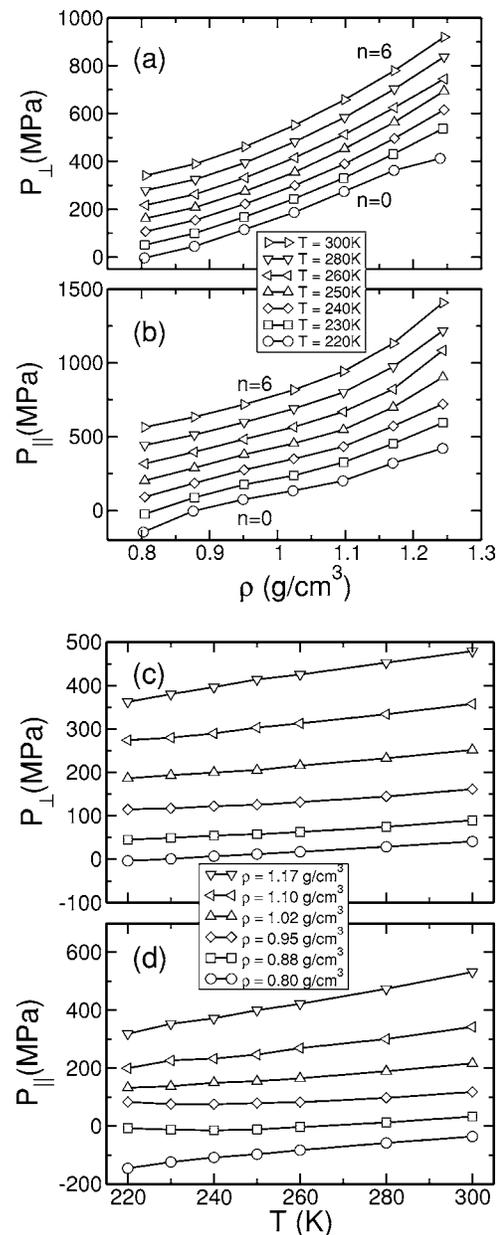


FIG. 3. (a) Transverse pressure P_{\perp} (perpendicular to the walls) and (b) lateral pressure P_{\parallel} (parallel to the walls) as functions of density ρ for all simulated T . For clarity, each curve is shifted by $n \times 50$ MPa for P_{\perp} for and $n \times 100$ MPa for P_{\parallel} . While P_{\perp} shows no inflection for all T , P_{\parallel} starts to flatten at $\rho \approx 0.87$ g/cm³ for the lowest T presented. This is consistent with the possibility of a second critical point at T lower than T_c of bulk water. (c) P_{\perp} and (d) P_{\parallel} as a function of temperature for different densities. Only P_{\parallel} displays minima, indicating the presence of a T_{MD} .

We also plot the T_{MD} for bulk water (from Ref. [27]) and confined water in Fig. 5. A +40 K temperature shift in the T_{MD} of confined water overlaps these loci. Thus, the effect of the hydrophobic walls in our system seems to be to shift the P - T phase diagram by $\Delta T \approx -40 \pm 5$ K with respect to bulk water. This is consistent with the second critical point shifting to lower T .

Figure 6 shows the calculated potential energy for the lowest simulated temperature $T = 220$ K. We note two

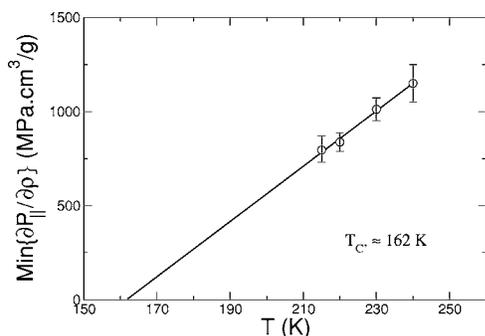


FIG. 4. Minimum values of $(\partial P_{||}/\partial \rho)_T$ as a function of temperature obtained from Fig. 3(d). The critical temperature T_C of the critical point of the LL transition is crudely estimated by extrapolation.

minima for the densities around $\rho=0.88 \text{ g/cm}^3$ and $\rho=1.39 \text{ g/cm}^3$, respectively. Since the free energy F is given by

$$F = K + U - TS, \quad (6)$$

where K , U , and S are the kinetic energy, potential energy and entropy respectively, at small T , an extremum in U suggests an extremum in F . Hence, the emergence of two minima at small T further supports the possibility of two stable liquids at low and high densities, respectively.

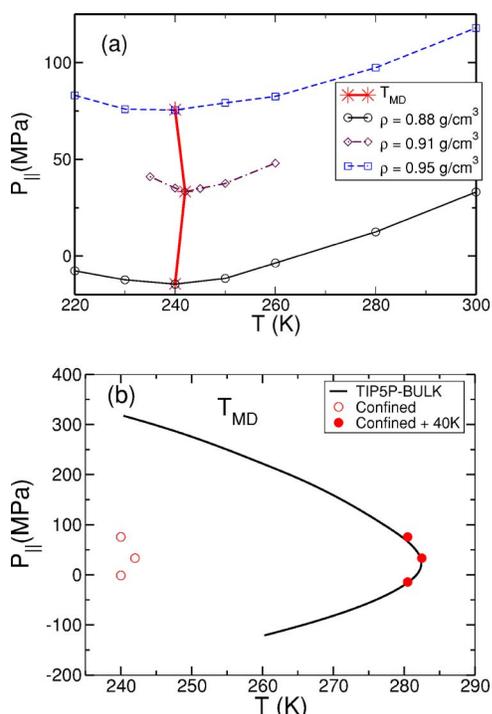


FIG. 5. (Color online) A comparison of T_{MD} line for bulk and confined water. (a) Isochores corresponding to the densities $\rho=0.88 \text{ g/cm}^3$, $\rho=0.91 \text{ g/cm}^3$, and $\rho=0.95 \text{ g/cm}^3$ which have a minimum. The locus of the points where $(\partial P_{||}/\partial T)_V=0$ denotes the T_{MD} (shown as asterisk). (b) If the T_{MD} for confined water (open circles) were shifted by 40 K (filled circles), it would overlap the T_{MD} for bulk water (solid line).

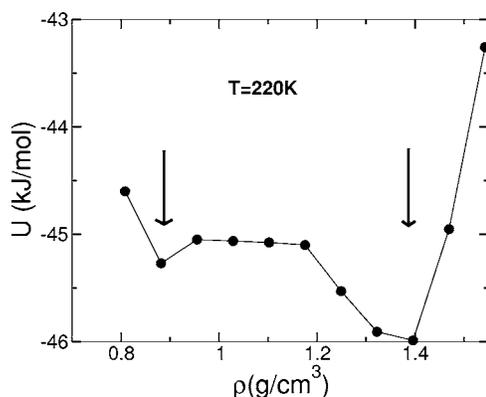


FIG. 6. Potential energy as a function of ρ for the lowest temperature we simulate, indicating the possibility of two stable liquid states at $\rho=0.88 \text{ g/cm}^3$ and $\rho=1.39 \text{ g/cm}^3$. The two minima are indicated by vertical arrows. Additional simulations were done at high densities for $T=220 \text{ K}$.

IV. STATIC STRUCTURE

A. Radial distribution function

In Sec. II, we studied the structure of water along the direction perpendicular to the walls. To aid in comparing the structural properties to those of bulk water, we next focus on the lateral oxygen-oxygen radial distribution function (RDF) defined by

$$g_{||}(r) \equiv \frac{1}{\rho^2 V} \sum_{i \neq j} \delta(r - r_{ij}) \left[\theta\left(|z_i - z_j| + \frac{\delta z}{2}\right) - \theta\left(|z_i - z_j| - \frac{\delta z}{2}\right) \right]. \quad (7)$$

Here V is the volume, r_{ij} is the distance parallel to the walls between molecules i and j , z_i is the z coordinate of the oxygen atom of molecule i , and $\delta(x)$ is the Dirac δ function. The Heaviside functions $\theta(x)$ restrict the sum to a pair of oxygen atoms of molecules located in the same slab of thickness $\delta z=0.1 \text{ nm}$. The physical interpretation of $g_{||}(r)$ is that $g_{||}(r)2\pi r dr \delta z$ is proportional to the probability of finding an oxygen atom in a slab of thickness δz at a distance r parallel to the walls from a randomly chosen oxygen atom. In a bulk liquid, this would be identical to $g(r)$, the standard RDF.

Figure 7(a) shows the effect on $g_{||}(r)$ of increasing ρ at low T . At low ρ , the bilayer liquid is characterized by a RDF that resembles the RDF of bulk water at $\rho=0.99 \text{ g/cm}^3$ [63], with maxima at $r \approx 0.28, 0.45, \text{ and } 0.67 \text{ nm}$. The well-defined maxima and minima in both $\rho(z)$ [Fig. 2(a)] and $g_{||}(r)$ [Fig. 7(a)] indicate that at low T and low ρ , the liquid is highly structured, with a structure parallel to the walls similar to corresponding bulk liquid water at low density [64]. As ρ increases, the liquid becomes less structured as indicated by the decreasing height of the second and third peaks of $g_{||}(r)$, and by the disappearance of the sublayers in $\rho(z)$ [Fig. 2(a)]. A comparison of $g_{||}(r)$ at $\rho=1.25 \text{ g/cm}^3$ to the $g(r)$ for bulk water at high density from Ref. [64] shows that the distributions are very different. However, we find that if we calculate $g_{||}(r)$ for only $-0.17 \text{ nm} \leq z \leq 0.17 \text{ nm}$ (which cor-

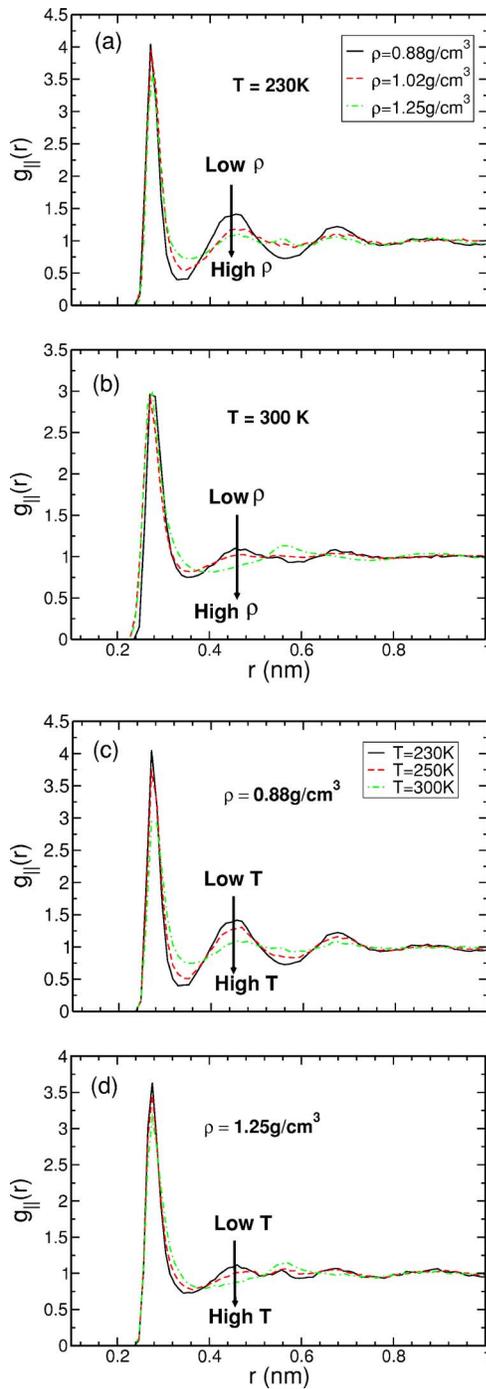


FIG. 7. (Color online) Lateral oxygen-oxygen radial distribution function $g_{\parallel}(r)$ for different densities at a constant temperature (a) $T=230$ K and (b) $T=300$ K; $g_{\parallel}(r)$ for different temperatures at a constant density (c) $\rho=0.88$ g/cm³ and (d) $\rho=1.25$ g/cm³.

responds to the location of the central layer), then $g_{\parallel}(r)$ at low ρ and high ρ resemble $g(r)$ for bulk water at both low and high densities, respectively. Thus the evolution of the structure parallel to the walls of the central layer mimics the structural changes when going from low-density bulk water to high-density bulk water.

The effect of increasing the density at $T=300$ K is shown in Fig. 7(b). The main effect of compression is to “redistrib-

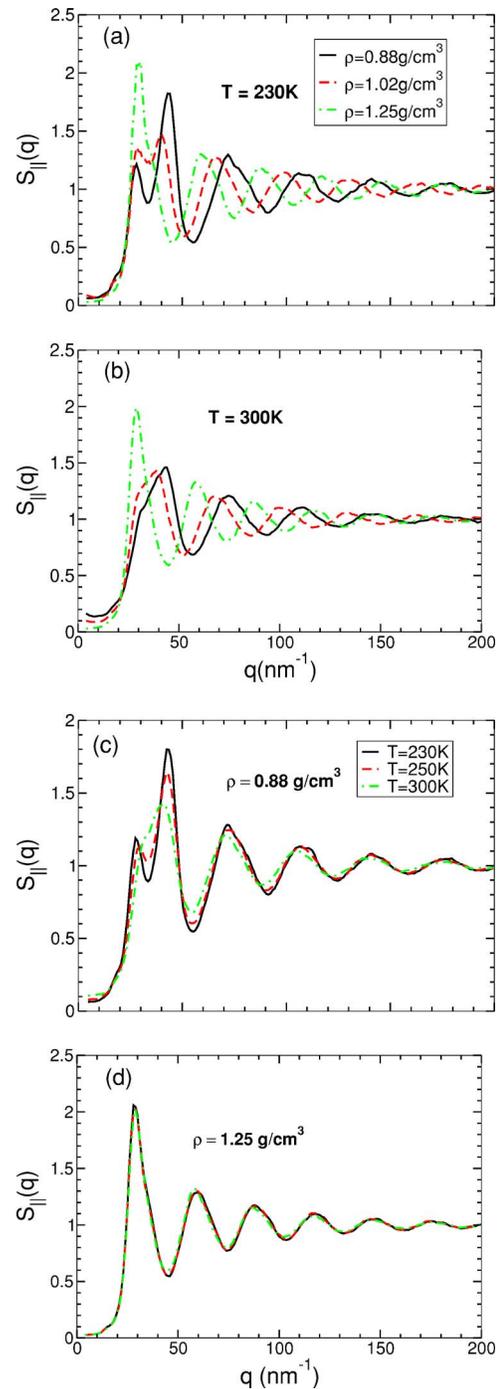


FIG. 8. (Color online) Lateral structure factor $S_{\parallel}(q)$ for different densities at a constant temperature (a) $T=230$ K and (b) $T=300$ K; $S_{\parallel}(q)$ for different temperatures a constant density (c) $\rho = 0.88$ g/cm³ and (d) $\rho = 1.25$ g/cm³.

ute” the molecules parallel to the walls. At $\rho=0.88$ g/cm³, $g_{\parallel}(r)$ is similar to the distribution shown in Fig. 7(a) at the same density, which resembles $g(r)$ for low-density bulk water. However, the maxima and minima in $g_{\parallel}(r)$ are less pronounced at $T=300$ K. At intermediate ρ , $g_{\parallel}(r)$ becomes very close to 1 for $r > 0.45$ nm. As density increases up to $\rho = 1.25$ g/cm³, a weak peak appears at $r \approx 0.9$ nm and a maximum occurs at 0.55 nm. Furthermore, the first maximum of

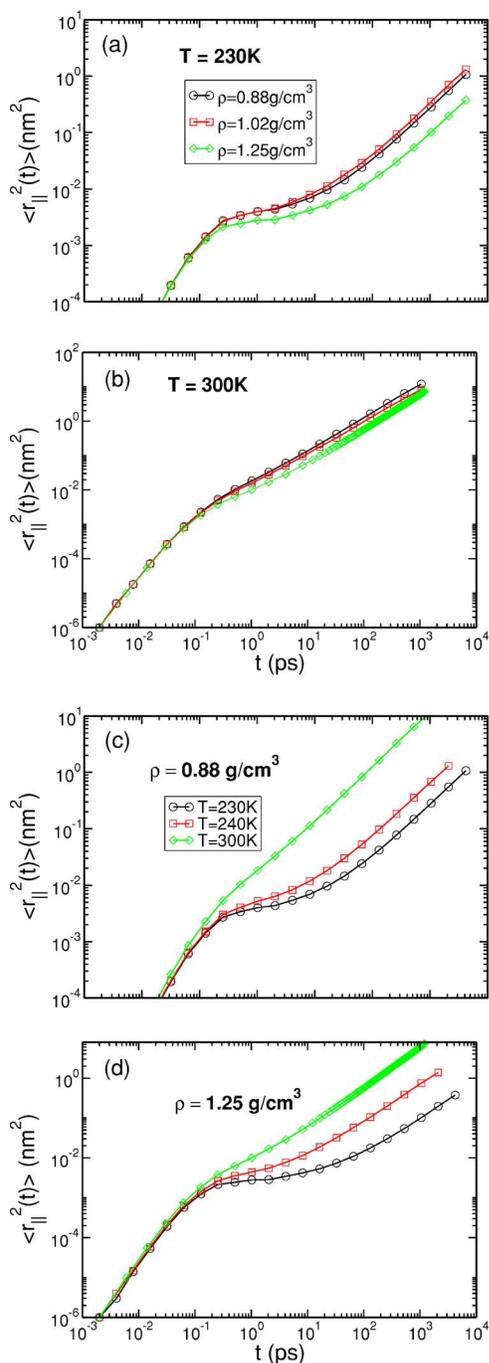


FIG. 9. (Color online) Lateral mean square displacement $\langle r_{\parallel}^2(t) \rangle$ for different densities at constant temperatures (a) $T=230\text{ K}$ and (b) $T=300\text{ K}$. $\langle r_{\parallel}^2(t) \rangle$ for different temperatures at constant densities (c) $\rho=0.88\text{ g/cm}^3$ and (d) $\rho=1.25\text{ g/cm}^3$.

$g_{\parallel}(r)$ becomes wider and the first minimum shift toward $r \approx 0.4\text{ nm}$. The resulting $g_{\parallel}(r)$ at high ρ and high T has many features of $g(r)$ for bulk water obtained experimentally at $T=268\text{ K}$ [64]. For example, the oxygen-oxygen $g(r)$ in Ref. [64] shows a weak peak at $r \approx 0.9\text{ nm}$, a clear maximum at 0.61 nm , and the first minimum is located at $r \approx 0.41\text{ nm}$. Furthermore, a shoulder in the first maximum of $g(r)$ develops at high density, which is consistent with the increase of the width of the first peak of $g_{\parallel}(r)$ in Fig. 7(b) as ρ increases.

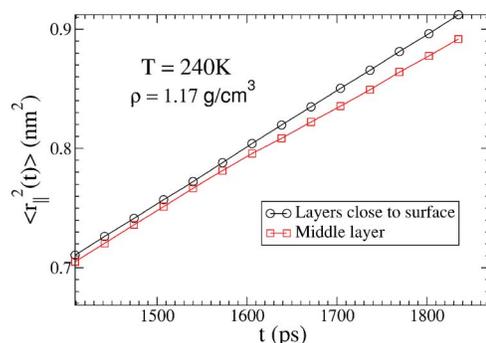


FIG. 10. (Color online) Diffusion in different layers, at a high enough density such that there are three layers. Shown is the lateral MSD in the two layers close to the surface and in the middle layer. We see that the water molecules close to the surface move faster than the water molecules in the middle layer.

To complete the comparison of the structure of confined water with the structure of bulk water, we also evaluate the T dependence of $g_{\parallel}(r)$. In Fig. 7(c), we show $g_{\parallel}(r)$ for various T at $\rho=0.88\text{ g/cm}^3$. The effects of increasing T at low ρ are similar to those observed in Fig. 7(a) when increasing ρ at low T . More specifically, the minima and maxima of $g_{\parallel}(r)$ become less pronounced as T increases and $g_{\parallel}(r)$ becomes flatter for $r > 0.5\text{ nm}$. We note that similar changes are also found in $\rho(z)$ when (i) increasing T at low ρ [Fig. 2(a)] and (ii) increasing ρ at low T [Fig. 2(c)].

Figure 7(d) shows the effect of T on $g_{\parallel}(r)$ when the density is fixed at $\rho=1.25\text{ g/cm}^3$. The effects of increasing T at high ρ [Fig. 7(d)] are similar to the effects of increasing ρ at a high T [Fig. 7(b)]. More specifically, the first minimum of $g_{\parallel}(r)$ shifts to a larger r and becomes less pronounced, whereas the second and third peaks located at 0.45 and 0.7 nm , respectively, merge and form an intermediate peak at $r = 0.55\text{ nm}$. This suggests that as ρ increases at high T , the preferred distance between second neighbors increases (parallel to the walls) and local tetrahedral order decreases. The emergence of a peak at 0.55 nm on heating at 1.25 g/cm^3 is in contrast to the behavior of bulk water, where the disappearance of the peaks at 0.45 and 0.7 nm gives rise to nearly featureless behavior of $g(r)$ beyond the first peak (see, e.g., Ref. [65]). Hence at high T , confinement gives rise to structure that would not be present in bulk systems, presumably because molecules orient relative to the walls. It is interesting to note that, as shown in Sec. II, there is almost no change in $\rho(z)$ with T , indicating that the rearrangement of the molecules parallel to the walls has no effect, on average, on the organization of molecules perpendicular to the walls.

B. Static structure factor

Next we calculate the lateral static structure factor $S_{\parallel}(q)$, defined as the Fourier transform of the lateral radial distribution function $g_{\parallel}(r)$,

$$S_{\parallel}(q) = \frac{1}{N} \sum_{i,j} \langle e^{i\mathbf{q} \cdot (\vec{r}_i - \vec{r}_j)} \rangle, \quad (8)$$

where the \mathbf{q} vector is the inverse space vector in the xy plane and r is the projection of the position vector on the xy plane.

The structure factor will be particularly useful for comparison with the crystal structure in Sec. VI, where distinct Bragg peaks in $S_{\parallel}(q)$ appear.

Figures 8(a) and 8(b) show the effect of density on the lateral structure factor $S_{\parallel}(q)$ for $T=230$ K and $T=300$ K. The structure factor of low-density and low-temperature confined water is similar to bulk water. The presence of a “prepeak” in $S_{\parallel}(q)$ at low ρ can be attributed to the existence of pronounced tetrahedral order in the low-temperature liquid. At high ρ or high T , this feature is reduced, as the molecular order becomes less tetrahedral and core repulsion dominates. This behavior is similar to bulk water, but with a shift to lower densities. We show the evolution of $S_{\parallel}(q)$ as a function of T for two different densities $\rho=0.88$ g/cm³ and $\rho=1.25$ g/cm³ in Figs. 8(c) and 8(d). The structure of low-temperature water for $\rho=0.88$ g/cm³ is similar to high density bulk water. When the temperature is increased, the repulsive region of the potential begins to dominate and tetrahedrality is reduced. The first two peaks in the structure factor merge to form a single peak [Fig. 8(c)]. However, at high density $\rho=1.25$ g/cm³, a change in temperature does not change the structure factor significantly [Fig. 8(d)]. Similar behavior is seen in bulk water [66].

V. DYNAMICS

Thus far, we have seen that if a LL transition exists for confined water, it is shifted to lower T than for bulk water, and that the tetrahedral order that gives rise to this behavior is also suppressed. Hence, it is natural to consider whether the dynamic properties of confined water exhibit the same temperature shift found for the thermodynamic properties relative to bulk water. For example, how is the maximum in diffusivity under pressure shifted under confinement? To compare with the bulk system, we calculate the lateral mean square displacement (MSD). We can evaluate the diffusion coefficient D from the asymptotic behavior of the lateral MSD using the Einstein relation

$$\langle r_{\parallel}^2 \rangle = 2dDt, \quad (9)$$

where $\langle r_{\parallel}^2 \rangle$ is the mean square displacement parallel to the walls over a sufficiently large time interval t , and d is the system dimension [68]. Since we calculate the diffusion only in the lateral directions, $d=2$.

Figures 9(a) and 9(b) show the dependence of the lateral MSD on ρ at fixed $T=230$ K and $T=300$ K. We also plot the dependence of the lateral MSD on T for two different temperatures in Figs. 9(c) and 9(d), using a log-log scale to emphasize the different mechanisms seen on different time scales: (i) an initial ballistic motion, where the lateral MSD is a quadratic function of time, $\langle r_{\parallel}^2(t) \rangle \sim t^2$; (ii) an intermediate “flattening” of the lateral MSD, due to the transient caging of molecules by their hydrogen bonded neighbors (This effect is most noticeable at the lowest T studied and does not occur at high T); and (iii) long time scales on which particles

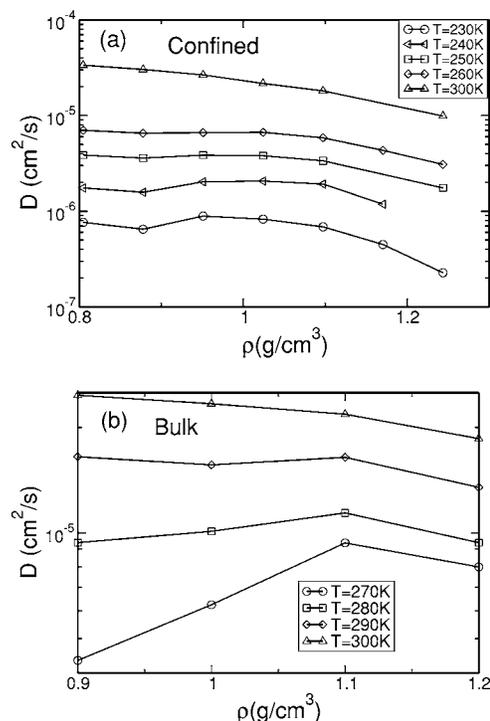


FIG. 11. Diffusion constant D as a function of density for various temperatures in (a) confined water and (b) bulk water. Confined water shows diffusion anomalies at lower temperatures than in bulk water. The first diffusion anomaly (the increase of diffusion with increase in density) in the confined system begins to appear when the temperature drops to $T \approx 250$ K, which is ≈ 40 K lower than in bulk water (which displays diffusion anomalies at $T \approx 290$ K).

diffuse randomly, and so $\langle r_{\parallel}^2(t) \rangle \sim t$.

A study of the diffusion of water molecules in different layers at high density (where there are three distinct layers) shows that the water molecules in the layer close to the surface move faster than the water molecules in the middle layer (see Fig. 10). This behavior of diffusion of molecules close to the surface, and away from the surface is similar to what is found for simple liquids [12].

To determine whether there is an anomaly in the density dependence of D , we plot D along isotherms, in Fig. 11(a) for confined water and in Fig. 11(b) for bulk water. For $T \approx 250$ K, we find that D has a maximum at $\rho \approx 1.05$ g/cm³. In bulk water, a similar behavior is found, but at $T \approx 290$ K, 40 K higher than the confined system. Moreover, this shift in a dynamic anomaly is consistent with the shift of thermodynamic anomalies. Qualitatively, the maximum in D can be understood as a competition between weakening or breaking of hydrogen bonds under pressure [67] (which increases D) and increased packing (which reduces D).

We next study the effect of confinement on the rotational motion of water molecules. The rotational motion was analyzed by calculating the rotational autocorrelation time for all the state points. We compare these results to the rotational autocorrelation time for bulk water for few state points. The rotational autocorrelation function $C_1(t)$ is defined as

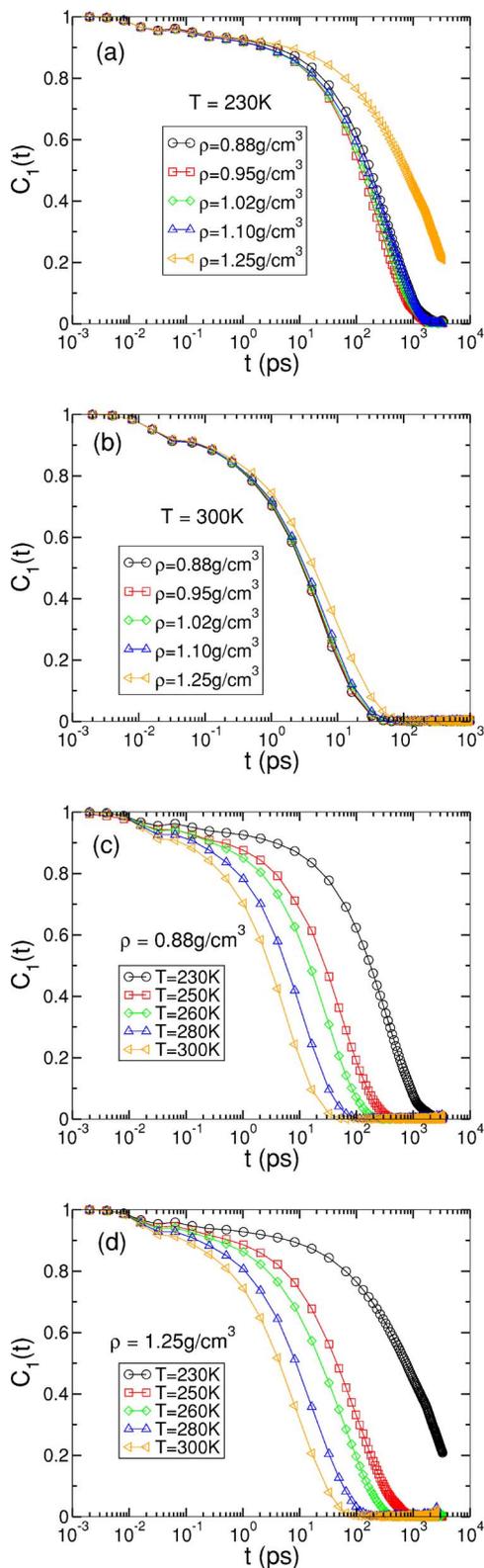


FIG. 12. (Color online) Dipole-dipole autocorrelation function $C_1(t)$ for different densities at constant temperature (a) $T=230$ K and (b) $T=300$ K. Also shown is $C_1(t)$ for different temperatures at constant density (c) $\rho=0.88$ g/cm³ and (d) $\rho=1.25$ g/cm³.

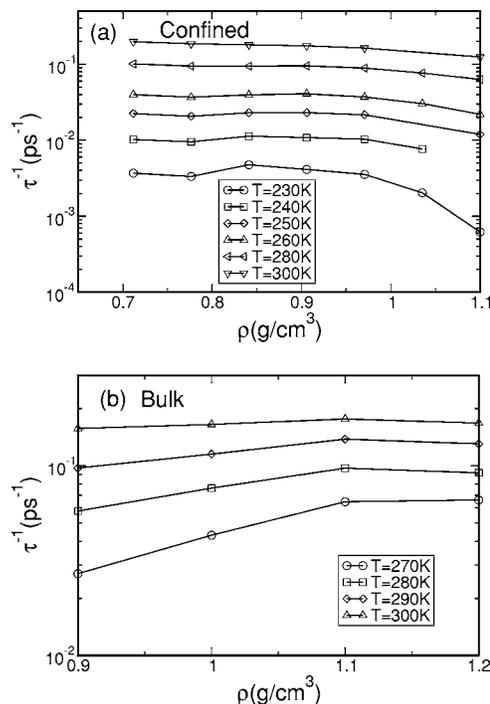


FIG. 13. Inverse orientational autocorrelation time as a function of density for (a) confined water and (b) bulk water. Inverse autocorrelation time for confined water follows a similar behavior as bulk water, but temperature shifted by ≈ 40 K to lower T .

$$C_1(t) \equiv \frac{1}{N} \left\langle \sum_{i=1}^N \vec{e}_i(t) \cdot \vec{e}_i(0) \right\rangle, \quad (10)$$

where $\vec{e}_i(t)$ is the unit dipole vector of molecule i at time t . For large times, $C_1(t)$ can be fit with a stretched exponential function

$$C_1(t) = A e^{(-t/\tau)^\beta}, \quad (11)$$

where A , $\beta \leq 1$ are constants and $\tau = \tau(\rho, T)$ is the orientational autocorrelation time, which depends on both density and temperature. In Fig. 12, we show $C_1(t)$ for different temperatures and densities. Figure 12(a) shows the inverse of the orientational autocorrelation time which is proportional to the rotational diffusion. For a comparison with the bulk water rotational diffusion, τ^{-1} is also shown in Fig. 12(b). Both the translational diffusion constant [Fig. 11(a)] and the inverse of the orientational autocorrelation time [Fig. 13(a)] show similar behavior. Similar results have been found for bulk water with the SPC/E potential [68]. For low temperature, the maxima occur at the same density for D and τ^{-1} . However at high temperatures ($T \geq 250$ K), where D is a monotonically decreasing function of density, τ^{-1} has a maximum and a minimum similar to its bulk counterpart [Fig. 13(b)].

VI. CRYSTALLIZATION OF TIP5P CONFINED WATER

Bulk TIP5P water crystallizes within the simulation time at high densities and for densities 1.15 g/cm³ and 1.20 g/cm³

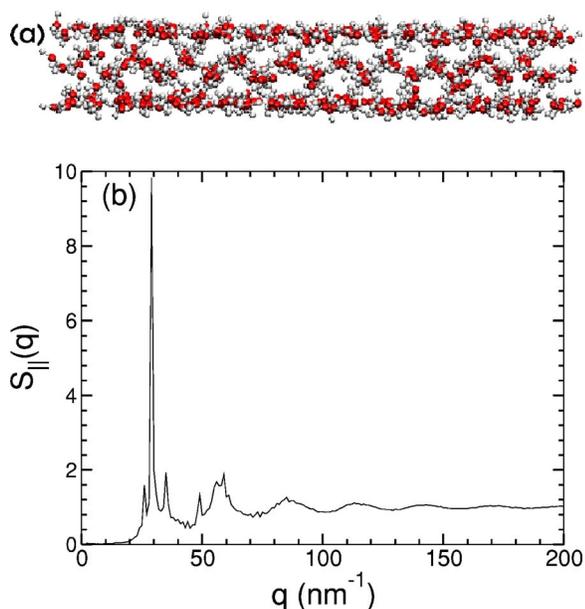


FIG. 14. (Color online) Trilayer ice with $L_z=1.1$ nm and $\rho=1.32$ g/cm³ formed when the system is cooled from $T=320$ K to $T=260$ K. (a) A lateral snapshot. (b) Lateral structure factor.

the crystallization time has a minimum at $T=240$ K [27]. We next investigate whether crystallization occurs in confinement, and whether the structure differs because of the surface effects. It has been found experimentally that water confined in hydrophobic carbon nanopores does not crystallize, even at very low temperatures [13]. However, the crystallization of confined water is seen in some simulations [48,49]. We find that our system crystallizes to what appears to be a trilayered ice structure at high density and that the resultant ice has a density $\rho=1.32$ g/cm³. A similar crystallization appears in simulations when an electric field is applied in a lateral direction [50]. The surface in this kind of simulation exhibits the embedded crystal structure of silica, where the oxygen and silicon atoms are arranged in out-of-registry order. This suggests that the crystalline form we find in confinement is weakly dependent on the morphology of the surface. We show the structure of the ice and the static lateral structure factor in Fig. 14.

To investigate whether confined water crystallizes in the different crystalline form when the separation between the plates is different, we repeat our simulation for a plate separation of 0.7 nm, and again the system crystallizes at $T=260$ K. The system crystallizes into a monolayer ice, also seen in simulations in Ref. [49]. The ice structure and its lateral static structure factor are shown in Fig. 15. The den-

TABLE I. Thermodynamic properties of the crystals formed at $T=260$ K

Plate separation (nm)	Pressure $P_{(\text{MPa})}$	Potential energy (kJ/mol)
0.7	-90.30	-39.86
1.1	652.89	-44.08

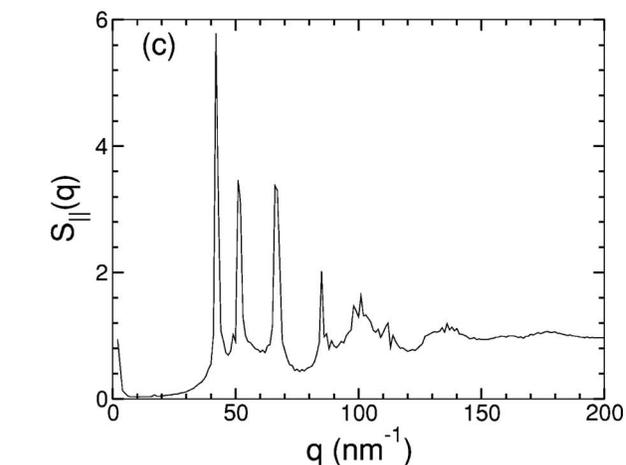
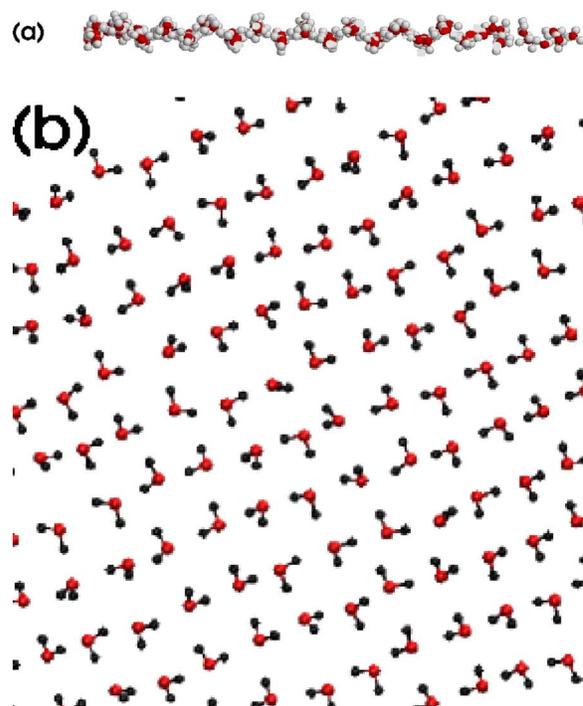


FIG. 15. (Color online) Monolayer ice with $L_z=0.7$ nm and $\rho=0.93$ g/cm³ spontaneously formed at $T=260$ K. (a) A lateral view (b) top view, and (c) lateral structure factor.

sity of monolayer ice is $\rho=0.93$ g/cm³. We list the temperature, pressure, and potential energy for these crystals in Table I.

VII. CONCLUSIONS

We have systematically investigated the effect of confinement on TIP5P water between two parallel smooth hydrophobic plates, separated by 1.10 nm. We found that the overall phase diagram is shifted to lower temperature and lower density compared to bulk TIP5P water by $\Delta T \approx -40$ K. The shift to lower temperature compared to the bulk water can be qualitatively understood. Since the confinement walls do not form any hydrogen bonds with water molecules the average number of hydrogen bonds per molecule in confined system

is smaller than in bulk water. This is analogous to having bulk water at high temperatures.

We do not see a LL phase transition for the state points we have been able to simulate, but we do see a pronounced inflection in the P - ρ isotherms (Fig. 3), which is consistent with a LL phase transition at lower T . Since the phase diagram is shifted ≈ 40 K lower in temperature, our results are consistent with the possibility that there indeed is a LL transition at a temperature too low to simulate. This shift of thermodynamics qualitatively agrees with the theoretical predictions of Ref. [47]. The structure of confined water is similar to the structure of bulk water at a lower density and shows a similar evolution of the structure with changes in density and temperature. At a given temperature, as the density increases, water changes from a bilayer liquid (at low density) to a trilayer liquid (at high density). We find that the confinement affects the translational diffusion as well as the rotational motion of water molecules. The rotational diffu-

sion anomaly precedes the translational diffusion anomaly, just as occurs for bulk water.

We were able to crystallize water for a few state points. It crystallizes spontaneously to a trilayer ice at $T=260$ K. Monolayer ice was formed when the separation between the plates was decreased to 0.7 nm. The crystalline structures are different from the polymorphs of bulk water and should be relevant for confined water.

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